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Physical Chemistry Keith J Laidler John H Meiser

The authors present physical chemistry in the historical manner it was discovered. While this is a traditional approach, it is not necessarily the most conducive to student's understanding of the material.

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Physical Chemistry CD Laidler Keith James Meiser John

CD for introductory physical chemistry. Abstract: This book's first edition appeared in 1972 and has succeeded in the difficult task of staying current despite enormous changes and progress enjoyed by inorganic chemistry over the last 21 years.

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Arrhenius Equation sas upenn edu

1 Arrhenius Equation The rate constant k for a reaction is related to the absolute temperature T by the equation: $k = Ae^{-B/T}$ A and B are constants and this equation can be also as $k = Ae^{-E/RT}$

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David W Ball, Essentials of Physical Chemistry , Thomson 2014 tel 22900318 ext232 or 0937424199 Laidler, Meiser, Sanctuary, Physical Chemistry 4th ed., Houghton Mifflin Co., 2003 tel 89121188

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Chapter 4 Physical Chemistry Laidler Meiser Sanctuary

This clip from the Physical Chemistry book explains how concentration changes with pressure. Changing the pressure on a gas can change the number of moles. This leads us to the concept of the chemical potential: each compound has a potential to react and has its own molar Gibbs free energy. The greater the chemical potential is negative (at constant T and P) the greater the potential to react and minimize the Gibbs energy for the whole reaction.

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